BATCH DRYING OF ACETIC ACID BY AZEOTROPIC DISTILATION (BATCH 3)

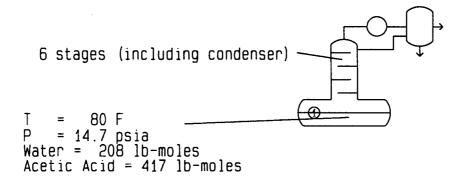
Water and acetic acid are difficult to separate. Production of dry acetic as a bottoms product while taking water overhead with a low acetic acid content is not easy in conventional distillation; many trays are required, and the column must be operated at a high reflux.

In this example, acetic acid containing 15 wt % water is dehydrated using ethyl acetate as an entrainer. Ethyl acetate and water are only partially miscible, thus they form a heterogeneous azeotrope. Ethyl acetate and water are easy to separate. The procedure is as follows: Bring the column to total reflux. Decant the water layer, while refluxing the ethyl acetate layer. The refluxing ethyl acetate effectively washes down to any acetic acid, so that the overhead product contains little acetic acid. As water is removed, the rate at which water is decanted drops off. At the point of diminishing returns, the decanter is taken out of service and ethyl acetate (containing some water and acetic acid) is taken as distillate to a different accumulator. When the ethyl acetate content of material in the pot has dropped to a sufficiently low concentration, the distillation is discontinued, and the pot is dumped out. Wet acetic acid for the next run is charged to the pot, and the ethyl acetate (containing some acetic acid and a small amount of water) is recycled.

In batch distillation, it is often difficult to decide on the details of a run at the start. Accordingly, the distillation is carried out in four steps. After each step, the user can stop to decide what to do in the next step; indeed, even the number of steps does not have to be fixed in advance. This strategy is used rather than planning details of the procedure at the start.

- Step 1 Decant water until the water rate from the decanter drops to 35% of the original rate. This is best done by plotting the water rate (and contained ethyl acetate) as a function of time, then pressing ESC.
- Step 2 The decanter is not used; instead the overhead is run in a conventional manner to remove part of the ethyl acetate and a large part of the remaining water. A 1/1 reflux ratio is used. The mass fufraction water and acetic acid in the distillate are plotted. When the acetic acid content of the distillate reaches 5.0 wt%, this step is discontinued. (An alternate strategy would be to use the Stop Option to discontinue the operation when the acetic acid content of the distillate reaches 5.00 wt%).
- Step 3 The acetic acid content of the overhead product is set at 5.30 wt% and the required reflux ratio plotted. When the reflux ratio reaches 5/1, this step is continued. Note that the purity is 5.0 wt% at the start of this step; thus, stability is no problem. If step 2 were eliminated (go directly from Step 1 to Step 3), stability and convergence problems would be likely. The required reflux ratio is followed on the screen and this step is discontinued when it reaches 5/1.
- **Step 4** Fix reflux ratio at 5/1 and plot the ethyl acetate content of the pot residue. When it reaches 0.5 wt%, stop the distillation.

Azeotropic Distillation Separation of Acetic Acid and Water using Ethyl Acetate as an entrainer



Batch Charge:
Temperature= 80 F
Pressure= 14.7 psia
Ethyl Acetate = 284 Ibmoles

Water = 208 lbmoles

Acetic Acid = 417 Ibmoles

Column:

6 stages

4 Operating Steps

Total condenser with decant

Cond P = 14.7

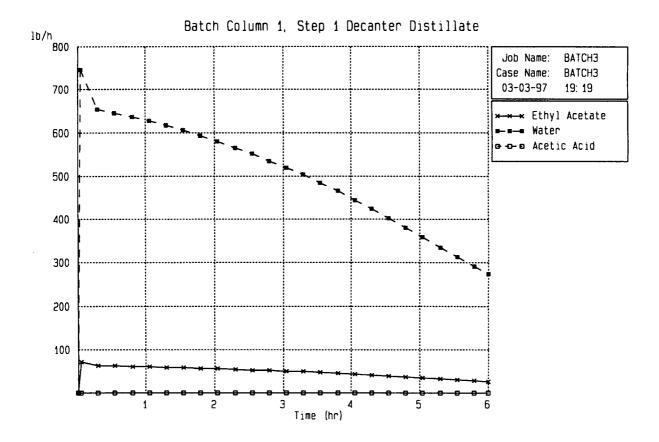
Volume holdup units

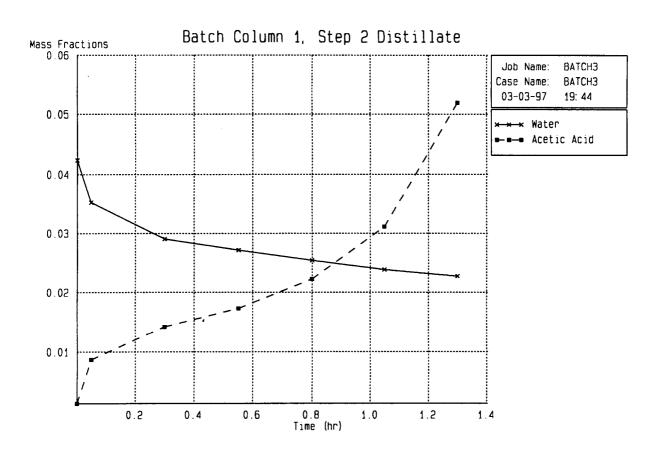
Condenser holdUp = 1.35

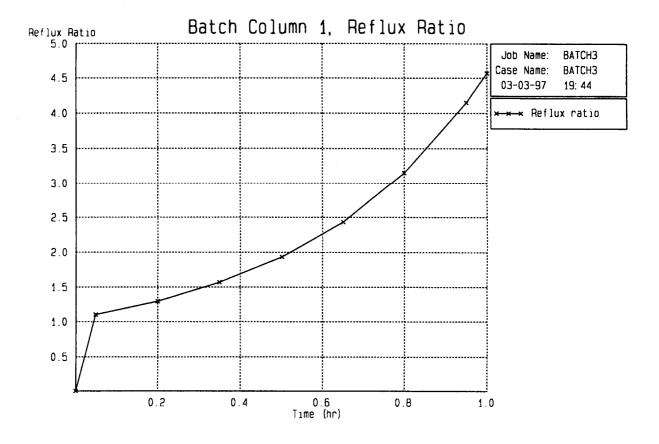
Stage holdup = 0.94

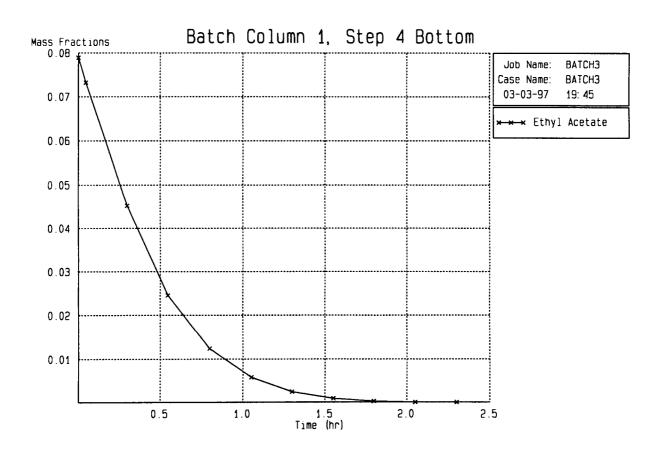
Operating Steps:

	1	2	3	4
Overhead Tank	1	2	3	4
Decant Tank	11	12	13	
Mode #1	Cond. Q	RR	Dist. Moss.Fr	RR
Value #1	-5	1	.05	5
Comp #1			3	
Mode #2	Dist.Mol.Flow	Reb. Q	Reb. Q	Reb. Q
Value #2	.00303	5	5	5
Comp #2				
Sub. Cond. T.	100	100	100	100
Stop Mode	Time	Time	Time	Time
Stop Value	6	1.3	1	2.3
Decant	on	off	off	off









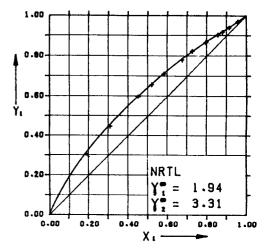
(1) WATER H20
(2) ACFTIC ACID C2H402

+++++ ANTOINE CONSTANTS REGION +++++
(1) 8.07131 1730.630 233.426 1- 100 0
(2) 8.02100 1936.010 258.451 18- 115 0

PRESSURE = 760.00 MM HG (1.013 BAR)

LIT: BRUSSET H., KAISER L., HOCQUEL J., CHIM. IND., GENIE CHIM. 99, 237 (1968).

CONSTANTS:	112	A21	ALPHA12
MARGULES VAN LAAR WILSCN NRTL	C.4178 0.4973 705.5876 712.1791	0.9533 1.6623 111.6579 330.1059	1.4032
. MI QUAC	-55.4002	116.7514	



EXPERIMENTAL DATA	MARGULES	VAN LAAR	WILSON	NR T L	UNIQUAC
T PEG C X1 Y1	DIFF T DIFF YI	DIFF T DIFF YI	DIFF T DIFF Y1	DIFF T DIFF YI	DIFF T DIFF YI
118.30	0.11	0.11 0.0 -0.61 -0.0070 -0.30 -0.0224 -0.18 -0.0181 -0.04 -0.0172 -0.03 -0.0124 0.25 -0.0077 0.00 0.0039 -0.06 0.0042 -0.23 0.0029 -0.18 0.0020 -0.24 0.0006 -0.21 -0.0011 0.00 0.0	0.11 0.0 -0.71 -0.0048 -0.39 -0.020C -0.24 -0.0154 -0.08 -0.0147 -0.07 -0.0102 0.23 -0.0063 -0.01 0.0048 -0.05 0.0043 -0.22 0.0029 -0.17 0.0020 -0.23 -0.0069 -0.20 -0.0009 0.00 0.0	0.11 0.0 0.10 -0.0077 0.02 -0.0089 -0.20 -0.0000 -0.13 -0.0021 -0.14 -0.018 0.19 -0.0050 -0.00 0.0028 0.01 0.0002 -0.13 -0.0007 -0.06 -0.0007 -0.12 -0.9002 -0.13 0.0002 -0.13 0.0002	0.11
MEAN DEVIATION:	3.24 3.0107	0.19 0.0083	0.22 0.0073	0.10 0.0025	0.19 0.0078
MAX. DEVIATION:	0.76 J.u316	C.61 G.0224	0.71 0.0200	0.20 0.0089	0.51 0.0220

CONSISTENCY

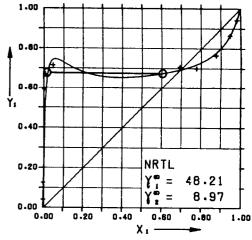
(1) ETHYL ACETATE	C4H8O2
(2) WATER	H2C

+++++ ANTOINE CONSTANTS REGION +++++
(1) 7.10179 1244.951 217.881 16- 76 C PETHCD 1 (2) 8.07131 1730.630 233.426 1- 100 C METHCD 2 +

PRESSURE= 760.00 MM HG (1.013 BAR)

LIT: ELLIS S.R.M., GARBETT R.D., IND.ENG. CHEM. 52,385(1960).

CONSTANTS:	A12	A21	ALPHA12
MARGULES	3.7016	1.8656	
VAN LAAR	3.8245	1.9685	
WILSON	1147.4297	2084.0256	
NRTL	935.6880	2316.3631	C.4104
UNIQUAC	757.5853	151.6420	



	ERIMENTAL	CATA	MAF	CULES	VAV	LAAR	WIL	.SON		IR TL	4.4.	01110
T DEG (: X1	¥1	DIFF T	DIFF Y1	DIFF T	DIFF Y1	DIFF T	DIFF Y1	DIFF T	DIFF Y1	DIFF T	QUAC DIFF Y1
98.85 96.20 86.50 76.20 70.55 70.45 71.45 73.35	0.0006 0.0011 0.0049 0.0086 0.0459 0.6900 0.7750 0.8737 0.9444 0.9807	0.0405 0.1256 0.5910 0.6680 0.7140 0.7070 0.6700 0.7650 0.8650 0.9411	0.17 -1.43 -4.33 -9.54 5.38 0.40 0.11 -0.74 -1.18	-0.0064 0.0426 0.3009 0.2495 -0.0466 5.1013 0.0527 0.0170 -0.0039 -0.0005	0.34 -1.14 -3.44 -8.44 5.05 -0.24 -0.66 -1.02 -1.09	-0.0121 0.0330 0.2774 0.2249 -0.0423 0.0676 0.0210 0.0021 -0.0033 -0.0063	C.46 -0.95 -3.48 -9.39 -4.33 -1.26 -1.26 -0.90 -0.50	-0.0164 0.0267 0.2785 0.2468 0.0841 -0.0013 -0.0305 -0.0129 0.0088 0.0060	0.43 -0.96 -2.44 -6.79 3.94 -1.13 -1.19 -0.96 -0.66	-0.0153 0.0270 0.2516 0.1899 -0.0277 0.016C -0.0191 -0.0101 0.0055 0.0028	0.33 -1.15 -3.27 -8.00 5.74 -0.33 -0.58 -0.81 -0.89	-0.0116 0.0333 0.2728 0.2155 -0.0490 0.0557 0.0146 0.0037 0.0015
	MEAN DEV	IATION:	2.41	6.083.	8.20	0.0530	2.25	·C.0712	9€	0.0565	2.15	0.0661
	MAX. DEV	:HOITA	9.54	0.3009	8.44	0.2774	9.39	0.2785	6.79	0.2516	8.00	0.2728

(1) ETHYL ACETATE C4H302

(2) ACETIC ACID C2H4O2

+++++ ANTOINE CCNSTANTS REGION +++++ CONSISTENCY
(1) 7.10179 1244.950 217.881 16- 76 C METHOD 1 +
(2) 7.55960 1644.048 233.524 17- 118 C METHOD 2 -

PRESSURE= 760.00 MM HG (1.013 BAR)

LIT: GARNER F.H., ELLIS S.R.M., PEARCE C.J. CHEM.ENG.SCI. 3,48(1954).

CONSTANTS: A12 A21 ALPHA12

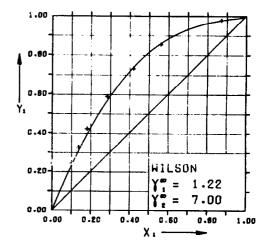
MARGULES Ø.2051 Ø.3976

VAN LAAR Ø.2313 Ø.4282

WILSON -997.0935 2738.1344

NRTL 844.2996 -436.9443

UNIQUAC 677.2519 -376.3130



EXPERIMENTAL DATA	MARGULES	VAN LAAR	WILSON	NRTL	UNICUAC
T DEG C X1 Y1	DIFF T DIFF Y1	DIFF T DIFF Y1	DIFF T DIFF Y1	DIFF T DIFF YI	DIFF T DIFF Y1
108.10 0.1380 0.3270 105.10 0.1820 0.4220 99.70 0.2860 0.5890 93.60 0.4200 0.7320 38.30 0.5610 0.8540 79.50 0.8720 0.9780	0.08 0.0125 -0.21 0.0232 0.02 0.0229 -0.19 0.0028 -0.56 0.0170 -0.89 0.0028	0.18 0.0121 -0.15 0.0240 -0.01 0.0255 -0.26 0.0103 -0.57 0.0169 -0.84 0.0027	-0.09 0.0136 -0.39 0.0228 0.02 0.0149 0.40 -0.0117 0.90 -0.0087 0.85 0.0003	0.01 0.0142 -0.31 0.0253 -0.11 0.0247 -0.25 0.0078 -0.49 0.0148 -3.81 0.0029	-0.02 0.0137 -0.33 0.0240 -0.06 0.0216 -0.16 0.0045 -0.40 0.0126 -0.79 0.0029
MEAN DEVIATION:	0.33 0.0144	0.34 0.0152	0.44 0.0120	0.33 0.0149	0.29 0.0132
MAX. DEVIATION:	0.89 0.0232	0.84 0.0255	0.90 0.0228	0.81 0.0253	0.79 0.0240